

Kinetic Molecular Theory Pogil 2005 Answers

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Kinetic Molecular Theory Pogil 2005

Kinetic Molecular Theory ©POGIL – 2005 5/5 Authored by Applications 1. There is a government warning on all aerosol cans that states: Do not store at a temperature above 120° F (50°C). a) Explain why this warning is required in terms of the relationship between temperature and pressure and the kinetic molecular theory.

Kinetic Molecular Theory - Ms. Lynch's Lessons

POGIL: Kinetic Molecular Theory. Modified from Foundations of Chemistryby David HansonPage 1 of 3. POGIL: Kinetic Molecular Theory. Learning Objectives. Identify the basic differences between particle behavior in the solid, liquid, and gaseous phases. Develop an understanding of the postulates of the kinetic molecular theory. Identify how temperature affects molecular motion. Apply the kinetic molecular theory to predict the outcome of everyday situations.

POGIL: Kinetic Molecular Theory

The kinetic molecular theory (KMT) is a simple microscopic model that effectively explains the gas laws described in previous modules of this chapter. This theory is based on the following five postulates described here. (Note: The term "molecule" will be used to refer to the individual chemical species that compose the gas, although some ...

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Answers Kinetic Molecular Theory Pogil

The Kinetic Molecular Theory Postulates The experimental observations about the behavior of gases discussed so far can be This theory is based on the following postulates, or assumptions. Gases are composed of a large number of particles that behave like hard, spherical

The Kinetic Molecular Theory - Purdue University

Information (Kinetic Molecular Theory) The observed behavior of gases, as expressed by the empirical gas laws, can be understood on the basis of the kinetic molecular theory, which developed over many years up to 1857, when Rudolf Clausius (1822-1888) published it in its most complete and elegant form.

Chem 116 POGIL Worksheet - Week 2 Gas Laws - Part 2

Kinetic Molecular Theory states that gas particles are in constant motion and exhibit perfectly elastic collisions. Kinetic Molecular Theory can be used to explain both Charles' and Boyle's Laws. The average kinetic energy of a collection of gas particles is directly proportional to absolute temperature only.

Kinetic Molecular Theory | Boundless Chemistry

The average kinetic energy of the gas particles is directly proportional to the Kelvin temperature of the gas. 3/3 ©POGIL –2005 Authored by Edited by Linda Padwa and David Hanson, Stony Brook University Kinetic Molecular Theory Key Questions 1. What causes a gas to exert pressure when confined in a container? of 2.

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The average kinetic energy of the gas particles is directly proportional to the Kelvin temperature of the gas. 3/3 - 2005 Authored by Edited by Linda Padwa and David Hanson, Stony Brook University Kinetic Molecular Theory Key Questions 1. What causes a gas to exert pressure when confined in a container? 2.

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9.5 The Kinetic-Molecular Theory - Chemistry

The Kinetic-Molecular Theory (The basic assumptions of kinetic-molecular theory are: {Postulate 1 zGases consist of discrete molecules that are relatively far apart. zGases have few intermolecular attractions. zThe volume of individual molecules is very small compared to the gas's volume. {Proof - Gases are easily compressible.

CHAPTER 12 GASES AND KINETIC-MOLECULAR THEORY

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Theory Pogil Answer 1)theory developed in the late 19thcentury to account for the behavior of the atoms and molecules that make up matter 2)based on the idea that particles in all forms of matter are always in motion and that this motion has

[EPUB] Molecular Geometry Pogil Answer

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The temperature of a gas depends on its average kinetic energy avg(1/2mv 2) = 3/2kT. In other words, the energy of an ideal gas is entirely kinetic. The amazing thing about the kinetic molecular theory is that it can be used to derive the ideal gas law.

Kinetic Molecular Theory: The Kinetic Molecular Theory ...

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